

Paper Code

2019 (A)

Roll No.

Number:

2481

INTERMEDIATE PART-I (11th CLASS)

CHEMISTRY PAPER-I (NEW SCHEME)

GROUP-I

TIME ALLOWED: 20 Minutes

OBJECTIVE

MAXIMUM MARKS: 17

Note: You have four choices for each objective type question as A, B, C and D. The choice which you think is correct, fill that bubble in front of that question number. Use marker or pen to fill the bubbles. Cutting or filling two or more bubbles will result in zero mark in that question. Attempt as many questions as given in objective type question paper and leave others blank. No credit will be awarded in case BUBBLES are not filled. Do not solve questions on this sheet of OBJECTIVE PAPER.

Q.No.1

- (1) The largest number of molecules are present in:
(A) 3.6g of H_2O (B) 4.8g of C_2H_5OH (C) 12.8g of CO (D) 5.4g of N_2O_5
- (2) One mole of SO_2 contains:
(A) 6.02×10^{23} atoms of Oxygen
(B) 18.1×10^{23} molecules of SO_2 (C) 6.02×10^{23} atoms of S (D) 4 gram atoms of SO_2
- (3) The comparative rates at which the solutes move in paper chromatography depend on:
(A) The size of paper (B) R_f values of solutes
(C) Temperature of the experiment (D) Size of the chromatographic tank used
- (4) The deviation of a gas from ideal behaviour is maximum at:
(A) $-10^\circ C$ and 5.0 atm (B) $-10^\circ C$ and 2.0 atm (C) $100^\circ C$ and 2.0 atm (D) $0^\circ C$ and 2.0 atm
- (5) The order of rate of diffusion of gases NH_3 , SO_2 , Cl_2 and CO_2 is:
(A) $NH_3 > SO_2 > Cl_2 > CO_2$ (B) $NH_3 > CO_2 > SO_2 > Cl_2$
(C) $Cl_2 > SO_2 > CO_2 > NH_3$ (D) $NH_3 > CO_2 > Cl_2 > SO_2$
- (6) Amorphous solids:
(A) Have sharp melting points (B) Undergo clean cleavage when cut with knife
(C) Have perfect arrangement of atoms (D) Can possess small regions of orderly arrangement
- (7) Diamond is a bad conductor because:
(A) It has a tight structure (B) It has a high density
(C) There are no free electrons present in the crystal of diamond to conduct electricity
(D) Is transparent to light
- (8) In the ground state of an atom, the electron is present:
(A) In the nucleus (B) In the second shell (C) Nearest to the nucleus (D) Farthest from the nucleus
- (9) Quantum number values for 2P orbitals are:
(A) $n = 2$ $\ell = 1$ (B) $n = 1$ $\ell = 2$ (C) $n = 1$ $\ell = 0$ (D) $n = 2$ $\ell = 0$
- (10) The number of bonds in Nitrogen molecule is:
(A) One σ and one π (B) One σ and two π (C) Three sigma only (D) Two σ and one π
- (11) Which of the Hydrogen halides has the highest percentage of ionic character?
(A) HCl (B) HBr (C) HF (D) HI
- (12) The change in heat energy of a chemical reaction, at constant temperature and pressure is called:
(A) Enthalpy change (B) Heat of combustion (C) Bond energy (D) Internal energy change
- (13) The solubility product of $AgCl$ is $2.0 \times 10^{-10} \text{ mol}^2 \text{ dm}^{-6}$. The maximum concentration of Ag^+ ions in the solution is:
(A) $2.0 \times 10^{-10} \text{ mol dm}^{-3}$
(B) $1.41 \times 10^{-5} \text{ mol dm}^{-3}$ (C) $1.0 \times 10^{-10} \text{ mol dm}^{-3}$ (D) $4.0 \times 10^{-20} \text{ mol dm}^{-3}$
- (14) Which of the following solutions has the highest boiling point?
(A) 5.85 % solution of Sodium Chloride (B) 18.0% solution of Glucose
(C) 6.0 % solution of Urea (D) All have the same boiling point
- (15) 18g of Glucose is dissolved in 90g of water. The relative lowering of vapour pressure is equal to:
(A) $\frac{1}{5}$ (B) 5.1 (C) $\frac{1}{51}$ (D) 6
- (16) If the salt bridge is not used between the two half cells, then the voltage:
(A) Decreases rapidly (B) Decreases slowly (C) Does not change (D) Drops to zero
- (17) If the rate equation of a reaction $2A + B \rightarrow \text{products}$ is, $\text{rate} = K[A]^2[B]$, and A is present in large excess, then the order of reaction is: (A) 1 (B) 2 (C) 3 (D) None of these

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INTERMEDIATE PART-I (11th CLASS)

CHEMISTRY PAPER-I (NEW SCHEME)

GROUP-I

TIME ALLOWED: 2.40 Hours

SUBJECTIVE

MAXIMUM MARKS: 68

NOTE: - Write same question number and its part number on answer book, as given in the question paper.

SECTION-I

2. Attempt any eight parts. 8 × 2 = 16
- Why do the isotopes have same chemical but different physical properties?
 - Define gram formula. Give two examples.
 - What is Stoichiometry? Give its two assumptions.
 - Why is there a need to crystallize the crude product?
 - What do you mean by solvent extraction? Which law does control it?
 - How is absolute zero explained by drawing graph?
 - Calculate the value of gas constant "R" in S.I units.
 - Derive Avogadro's Law from Kinetic molecular theory of gases.
 - Lighter gases diffuse more rapidly than heavier gases. Give reasons.
 - One molal solution of urea is more dilute than one molar solution. Why?
 - Define Raoult's Law. Give one of its mathematical expression
 - What is discontinuous solubility curve? Give one example.
3. Attempt any eight parts. 8 × 2 = 16
- Why do fish and plants in ponds survive under blanket of ice during cold winters?
 - Define Polymorphism. Give one example.
 - Freshly cut metals show the property of metallic luster. Comment on the statement.
 - Write down any two properties of Molecular Solids.
 - Calculate the mass of electron from its e/m value.
 - Justify the statement that angular momentum of an electron revolving in orbit is quantized.
 - How was dual nature of electron verified by Davisson and Germer?
 - State Aufbau principle. Write electronic configuration of Sodium ($_{11}\text{Na}$) following this principle.
 - Define the given terms: (a) common ion effect (b) solubility product
 - What is the difference between heat and temperature?
 - Define pseudo first order reaction. Give one example.
 - Write down two examples to explain the activation of catalyst.
4. Attempt any six parts. 6 × 2 = 12
- Differentiate between exothermic and endothermic reactions.
 - Define Ionization energy. How does it vary in periodic table?
 - Ionic radii of anions are greater than their parent atoms. Why?
 - CO_2 is non-polar whereas H_2O is polar molecule. Give reason.
 - Define Dipole Moment. Give its various units.
 - State 1st law of thermodynamics.
 - Define oxidation number. Calculate oxidation number of 'Mn' in KMnO_4 .
 - Differentiate between a primary cell and a secondary cell.
 - Write electrochemical reactions taking place in Alkaline battery.

SECTION-II

NOTE: - Attempt any three questions.

- 5.(a) Describe combustion analysis to determine mass percentages of C, H and O in an organic compound. 4
- (b) What is meant by Hydrogen Bonding? How it explains the helix structure of proteins? 4
- 6.(a) Write defects in Bohr's model of an atom. 4
- (b) One mole of methane gas is maintain at 300K its volume is 250cm^3 . Calculate the pressure exerted by the gas. 4
- 7.(a) What is Electron Affinity? How does it show variation along groups and periods in the periodic table? 4
- (b) Describe the Hess's Law of Constant Heat Summation and give one example to explain it. 4
- 8.(a) $\text{N}_{2(g)}$ and $\text{H}_{2(g)}$ combine to give $\text{NH}_{3(g)}$. The value of K_c in this reaction at 500°C is 6.0×10^{-2} . Calculate the value of K_p for this reaction. 4
- (b) How does Arrhenius equation help us to calculate the energy of activation of a reaction? 4
- 9.(a) Write Landsberger's method for determination of elevation of boiling point. 4
- (b) Define voltaic of Galvanic cell. Write its function with chemical equations. 4