For a reaction NaOH + HCl	GROUP : FIRST	11" CLASS -	- 1° Annual 2024	MARKS:17	
you think is correct, fill that circle in front of that question number. Use marker or pen to fill the circles. Cutting or filling two or more circles will result in zero marks in that question.    Dall		OB	JECTIVE		
you think is correct, fill that circle in front of that question number. Use marker or pen to fill the circles. Cutting or filling two or more circles will result in zero marks in that question.    Dall -1 - 24	IOTE: You have for	ur choices for each obj	ective type question as A , B ,	C and D . The choice which	
For a reaction NaOH + HCl	vou think is	correct, fill that circle i	n front of that question numbe	r. Use marker or pen to fill	
For a reaction NaOH + HCI	the circles.	Cutting or filling two or r	nore circles will result in zero m	arks in that question.	
(A) Heat of neutralization (C) Heat of formation (D) Heat of combustion An excess of silver nitrate is added to barium chloride solution and precipitates removed by filtration , what are the main ions in the filtrate? (A) Ba²¹ and NO₃ only (B) Ag² , Ba²¹ and NO₃ only (C) Ag² and NO₃ only (D) Ba²¹ , NO₃ and .Cl Which of the following solution has the highest boiling point? (A) 18 % solution of glucose (B) 6.0 % solution of urea (C) 5.85 % solution of sodium chloride (D) All have the same boiling point (fa strip of Cu metal is placed in a solution of FeSO₃: (A) Cu will be deposited (C) Cu and Fe both dissolved (D) No reaction takes place With increase of 10 °C temperature the rate of reaction doubles. This increase in rate of reaction is due to: (A) Increase in number of effective collisions. (B) Increase in activation energy of reaction. (D) Decrease in activation energy of reaction. (D) Decrease in the number of collisions between reactant molecules. One mole of SO₂ contains: (A) 6.02 x 10²3 atoms of oxygen (C) 18.1 x 10²3 molecules of SO₂ How many particles are called fundamental particles of an atom? (A) 3 (B) 5 (C) 100 (D) 6 What are the units of Ry value? (A) Cm (B) Cm³ (C) (Dm³ (D) No units Which of the following cannot sublime? (A) Naphthalene (B) Iodiae (C) Ammonium chloride (D) MnO₂ If absolute temperature of a gas to doubled and the pressure is reduced to one half, the volume of the gas will (A) Be doubled (B) Refused 1/4 (C) Increases four times (D) Remain unchanged Partial pressure of oxygen in lurgs (in Torr) is: (A) 150 (B) 116 (C) 760 (D) 159 Molecules of CO₁ in dry (ce form the: (A) 150 (B) 116 (C) 760 (D) 159 Molecules of CO₁ in dry (ce form the: (A) 150 (B) 160 (C) 200 (D) 5 Molecules of CO₁ in dry (ce form the: (A) 150 (B) 160 (C) 200 (D) 5 Molecules of CO₁ in dry (ce form the: (A) 150 (B) 160 (C) 200 (D) 5 Molecules of CO₁ in dry (ce form the: (A) 150 (B) 160 (B) 16	QUESTION NO. 1		V -1		
(C) Heat of formation (D) Heat of combustion An excess of silver nitrate is added to barium chloride solution and precipitates removed by filtration , what are the main ions in the filtrate?  (A) Ba²* and NO₃ only (D) Ba²*, NO₃ and Cl Which of the following solution has the highest boiling point?  (A) 18 % solution of glucose (B) 6.0 % solution of urea  (C) 5.85 % solution of sodium chloride (D) Alt have the same boiling point if a strip of Cu metal is placed in a solution of FeSO₁:  (A) Cu will be deposited (B) Fe is precipitated out  (C) Cu and Fe both dissolved (B) Fe is precipitated out  (C) Cu and Fe both dissolved (D) No reaction takes place  With increase of 10 °C temperature the rate of reaction doubles. This increase in rate  of reaction is due to:  (A) Increase in number of effective collisions. (B) Increase in activation energy of reaction.  (D) Decrease in number of collisions between reactant molecules.  One mole of SO₂ contains:  (A) 6.02 x 10²3 atoms of oxygen  (B) 6.02 x 10²3 atoms of sulphur  (C) 18.1 x 10²3 molecules of SO₂  How many particles are called fundamental particles of an atom?  (A) 3 (B) 5 (C) 100 (D) 6  What are the units of Rr value?  (A) Cm (B) Cm² (C) dm² (D) No units  Which of the following cannot sublime?  (A) Naphthalene (B) Iodine (C) Ammonium chloride (D) MnO₂  If absolute temperature of a gas is doubled and the pressure is reduced to one half, the volume of the gas will  (A) Be doubled (B) Reduced 1/4 (C) Increases four times (D) Remain unchanged Partial pressure of oxygen inclures (In) rorr) is:  (A) 150 (B) 115 (E) 760 (D) 159  Molecules of CO₂ in dry ide form the:  (A) Molecular crystals (B) Noinic crystals (C) Covalent crystals (D) Any type of crystals vapour pressure is not officeted by:  (A) Temberature (B) Intermolecular forces (C) Surface area (D) Pressure  Wave number of the light-emitted by a certain source is 2 x 10 <sup>6</sup> m². The wavelength of this light will be:  (A) 7 The bond order of helium molecule is:  (A) 3 (B) 2 (C) B² (D) Zero	1 For a reaction NaOH			y is called :	
An excess of silver nitrate is added to barium chloride solution and precipitates removed by filtration, what are the main ions in the filtrate?  (A) Ba² and NO₃ only  (B) Ag², Ba²*, AO₃ and Cſ  Which of the following solution has the highest boiling point?  (A) 18 % solution of glucose  (B) 6.0 % solution of urea  (C) 5.85 % solution of sodium chloride  (C) Sun solution of sodium chloride  (B) Aff solution of urea  (C) Cu and Fe both dissolved  (C) Cu and Fe both dissolved  (B) Fe is precipitated out  (C) Cu and Fe both dissolved  (B) No reaction takes place  With increase of 10 °C temperature the rate of reaction doubles. This increase in rate of reaction is due to:  (A) Increase in aumber of effective collisions.  (B) Increase in activation energy of reaction.  (D) Decrease in athe number of collisions between reactant molecules.  One mole of SO₂ contains:  (B) 6.02 x 10²3 atoms of soxygen  (C) 18.1 x 10²3 molecules of SO₂  How many particles are called fundamental particles of an atom?  (A) 3 (B) 5 (C) 100  (b) 6  What are the units of Rr value?  (A) Cm (B) Cm² (C) dm² (D) No units  Which of the following cannot sublime?  (A) Naphthalene  (B) Iodine  (C) Ammonium chloride  (D) MnO₂  If absolute temperature of a gas is doubled and the pressure is reduced to one half, the volume of the gas will  (A) Be doubled  (B) Reduced 1/4 (C) Increases four times  (D) Remain unchanged  Partial pressure of oxygen in lungs (in) orry is:  (A) 150 (B) 16 (C) 760 (D) 159  Molecules of CO₂ in dry ice form the:  (A) Moleculer crystals  (B) Ionic crystals  (C) Covalent crystals  (D) Any type of crystals  Vapour pressure is not affected by:  (A) Temperature  (B) The primolecular forces  (C) Surface area  (D) Pressure  Wave number of the light emitted by a certain source is 2 x 10 <sup>6</sup> m². The wavelength of this light will be:  (A) On.m (B) 500 m (C) 200 n.m (D) 5 x 10 <sup>7</sup> m  Sadioactive copper emits:  (A) A - rays  (B) Fe is predictive to the pressure is not affected by:  (A) The bond order of helium molecule is:  (B) A - rays  (C)	(A) Heat of neutraliz	ation (B	Heat of reaction		
filtration , what are the main ions in the filtrate?  (A) Ba <sup>2+</sup> and NO <sub>3</sub> only  (B) Ag <sup>+</sup> , Ba <sup>2+</sup> and NO <sub>3</sub> only  (C) Ag <sup>+</sup> and NO <sub>3</sub> only  (D) Ba <sup>2+</sup> , NO <sub>3</sub> and Cl <sup>-</sup> Which of the following solution has the highest boiling point?  (A) 18 % solution of sodium chloride  (B) 6.0-% solution of urea  (C) 5.85 % solution of sodium chloride  (B) All have the same boiling point  (A) Cu will be deposited  (C) Cu and Fe both dissolved  (B) Fe is precipitated out  (C) Cu and Fe both dissolved  (C) Cu and Fe both dissolved  (B) Fe is precipitated out  (C) Cu and Fe both dissolved  (C) Uncrease of 10 °C temperature the rate of reaction doubles. This increase in rate  of reaction is due to:  (A) Increase in number of effective collisions.  (B) Increase in activation energy of reaction.  (D) Decrease in activation energy of reaction.  (D) Decrease in the number of collisions between reactant molecules.  One mole of SO <sub>2</sub> contains:  (A) 6.02 x 10 <sup>23</sup> atoms of oxygen  (C) 18.1 x 10 <sup>23</sup> molecules of SO <sub>2</sub> How many particles are called fundamental particles of an atom?  (A) 3  (B) 5  (C) 100  (D) 6  What are the units of R <sub>1</sub> value?  (A) Cm  (B) Cm <sup>3</sup> (C) dm <sup>3</sup> (D) No units  Which of the following cannot sublime?  (A) Naphthalene  (B) Iodine  (C) Ammonium chloride  (D) MnO <sub>2</sub> If absolute temperature of a gas is doubled and the pressure is reduced to one half , the volume of the gas will  (A) Be doubled  (B) Reduced 1/4  (C) Increases four times  (D) Remain unchanged  Partial pressure of oxygen in lungs (in orr) is:  (A) 150  (B) 116  (C) 760  (D) 159  Molecules of CO; in dry ige form tife:  (A) Molecules of CO; in dry ige form tife:  (A) Molecules of CO; in dry ige form tife:  (A) Molecules of CO; in dry ige form tife:  (A) Molecules of CO; in dry ige form tife:  (A) Temperature  (B) Intermolecular forces  (C) Surface area  (D) Pressure  4 Wave number of the light emitted by a certain source is 2 x 10 <sup>6</sup> m <sup>-1</sup> . The wavelength of this light will be:  (A) Cn (B) Gr (C) Br <sub>2</sub> (D) Jero  7 The bond order of helium molecule is	(C) Heat of formation				
(A) Ba²* and NO₃ only (B) Ag²*, NO₃ and Cl  Whitch of the following solution has the highest boiling point (C) Ag³ and NO₃ only (D) Ba²*, NO₃ and Cl  Whitch of the following solution has the highest boiling point (F)  (A) 18 % solution of glucose (B) .6.9 % solution of urea  (C) 5.85 % solution of sodium chloride (D) All have the same boiling point (G) Lou will be deposited (D) No reaction takes place (D) No reaction takes place (D) No reaction is due to:  (A) Cu will be deposited (D) No reaction doubles. This increase in rate of reaction is due to:  (A) Increase in number of effective collisions. (B) Increase in activation energy of reaction.  (D) Decrease in activation energy of reaction.  (D) Decrease in the number of collisions between reactant molecules.  CO emole of SO₂ contains:  (A) 6.02 x 10²³ atoms of oxygen (D) 4 gram atoms of SO₂  How many particles are called fundamental particles of an atom?  (A) 3 (B) 5 (C) 100 (D) 6  What are the units of Rf value?  (A) Cm (B) Cm³ (C) dm³ (D) No units  Which of the following cannot sublime.  (A) Naphthalene (B) Iodine (C) Ammonium chloride (D) MnO₂  If absolute temperature of a gas is doubled and the pressure is reduced to one half, the volume of the gas will  (A) Be doubled (B) Refused 1/4 (C) Increases four times (D) Remain unchanged  Partial pressure of oxygen in lungs (in torr) is:  (A) 150 (B) 116 (C) 760 (D) 159  Molecules of CO₃ in dry ide form tips:  (A) Molecular crystals (B) bonic crystals (C) Covalent crystals (D) Any type of crystals  Vapour pressure is not affected by:  (A) Temperature (B) Intermolecular forces (C) Surface area (D) Pressure  Wave number of the light-emitted by a certain source is 2 x 106 m³ . The wavelength of this light will be:  (A) 3 (B) CHCl³ (C) BF₃ (D) H₂O  The bond order of helium molecule is:  (A) A - rays (B) β - rays (C) γ - rays (D) Positive rays  Which of the following molecules have zero dipole moment?  (A) NH₃ (B) CHCl³ (C) BF₃ (D) H₂O	2 An excess of silver ni	trate is added to bariu	m chloride solution and precip	itates removed by	
(C) Ag* and NO <sub>3</sub> only  Which of the following solution has the highest boiling point?  (A) 18 % solution of glucose  (B) 6.0 % solution of urea  (C) 5.85 % solution of sodium chloride  (D) Alt have the same boiling point  if a strip of Cu metal is placed in a solution of FeSO <sub>4</sub> :  (A) Cu will be deposited  (B) Fe is precipitated out  (C) Cu and Fe both dissolved  (C) U and Fe both dissolved  (D) No reaction takes place  With increase of 10 °C temperature the rate of reaction doubles. This increase in rate  of reaction is due to:  (A) Increase in number of effective collisions.  (B) Increase in activation energy of reaction.  (D) Decrease in activation energy of reaction.  (D) Decrease in the number of collisions between reactant molecules.  One mole of SO <sub>2</sub> contains:  (A) 6.02 x 10 <sup>23</sup> atoms of oxygen  (D) 4 gram atoms of SO <sub>2</sub> How many particles are called fundamental particles of an atom?  (A) 3 (B) 5 (C) 100  (B) 6  What are the units of R <sub>1</sub> value?  (A) Cm (B) Cm <sup>3</sup> (C) dm <sup>3</sup> (D) No units  Which of the following cannot sublime?  (A) Naphthalene (B) lodge (C) Ammonium chloride (D) MnO <sub>2</sub> If absolute temperature of a gas is doubled and the pressure is reduced to one half , the volume of the gas will  (A) Be doubled (B) Reduced 1/4 (C) Increases four times (D) Remain unchanged Partial pressure of oxygen in lungs (in) orr) is:  (A) 150 (B) 116 (C) 760 (D) 159  Molecules of CO <sub>2</sub> in dry ice form the:  (A) 150 (B) 116 (C) 760 (D) 159  Molecules of CO <sub>2</sub> in dry ice form the:  (A) Molecules of CO <sub>3</sub> in dry ice form the:  (A) Temberature (B) Intermolecular forces (C) Surface area (D) Pressure  Wave number of the light emitted by a certain source is 2 x 10 <sup>6</sup> m <sup>-1</sup> . The wavelength of this light will be:  (A) 3 (B) CHCl <sub>3</sub> (C) BF <sub>3</sub> (D) H <sub>2</sub> O  7 The bond order of helium molecule is:  (A) 3 (B) CHCl <sub>3</sub> (C) BF <sub>3</sub> (D) H <sub>2</sub> O  7 The bond order of helium molecule is:  (A) 3 (B) CHCl <sub>3</sub> (C) BF <sub>3</sub> (D) T <sub>2</sub> O	filtration, what are t	he main ions in the filt	rate ?		
(C) Ag* and NO <sub>3</sub> only  (D) Ba²*, NO <sub>3</sub> and CI  Which of the following solution has the highest boiling point?  (A) 18% solution of glucose  (C) 5.85% solution of sodium chloride  (I) All have the same boiling point  If a strip of Cu metal is placed in a solution of FeSO <sub>4</sub> :  (A) Cu will be deposited  (B) Fe is precipitated out  (C) Cu and Fe both dissolved  (C) Cu and Fe both dissolved  (B) No reaction takes place  With increase of 10 °C temperature the rate of reaction doubles. This increase in rate of reaction is due to:  (A) Increase in number of effective collisions. (B) Increase in activation energy of reactants.  (C) Decrease in a the number of collisions between reactant molecules.  One mole of SO <sub>2</sub> contains:  (A) 6.02 x 10 <sup>23</sup> atoms of oxygen  (C) 18.1 x 10 <sup>23</sup> molecules of SO <sub>2</sub> How many particles are called fundamental particles of an atom?  (A) 3 (B) 5 (C) 100  (B) Agran atoms of SO <sub>2</sub> How many particles are called fundamental particles of an atom?  (A) Cm (B) Cm <sup>3</sup> (C) dm <sup>3</sup> Which of the following cannot sublime?  (A) Cm (B) Cm <sup>3</sup> (C) dm <sup>3</sup> Which of the gas will  (A) Be doubled (B) Reduced 1/4 (C) Increases four times (D) Remain unchanged Partial pressure of oxygen in lungs (in orr) is:  (A) 150 (B) 116 (G) 760 (D) 159  Molecules of CO; in dry ice form tife:  (A) Molecular crystals  (B) Intermolecular forces (C) Surface area (D) Pressure  (B) Intermolecular forces (C) Surface area (D) Pressure  (B) Intermolecular forces (C) Surface area (D) Pressure  (Wave number of the light emitted by a certain source is 2 x 10 <sup>6</sup> m <sup>1</sup> . The wavelength of this light will be:  (A) Temberature (B) Soom (C) 200 n.m (D) 5 x 10 <sup>7</sup> m  Radioactive copper emits:  (A) A - rays (B) β - rays (C) γ - rays (D) Positive rays  Which of the following molecules have zero dipole moment?  (A) NH <sub>3</sub> (B) CHCl <sub>3</sub> (C) BF <sub>3</sub> (D) H <sub>2</sub> O  The bond order of helium molecule is:  (A) 3 (B) 2 (C) 1 (D) Zero	(A) Ba <sup>2+</sup> and NO <sub>3</sub> o	only (I	B) $Ag^+$ , $Ba^{2+}$ and $NO_3^-$ only		
Which of the following solution has the highest boiling point?  (A) 18 % solution of glucose (B) 6.0 % solution of urea (C) 5.85 % solution of sodium chloride (D) Alt have the same boiling point if a strip of Cu metal is placed in a solution of FeSO <sub>4</sub> :  (A) Cu will be deposited (B) No reaction takes place with increase of 10 °C temperature the rate of reaction doubles. This increase in rate of reaction is due to:  (A) Increase in number of effective collisions. (B) Increase in activation energy of reactants. (C) Decrease in activation energy of reaction. (D) Decrease in the number of collisions between reactant molecules. (D) Decrease in the number of collisions between reactant molecules. (D) Decrease in the number of sollisions between reactant molecules. (D) 4 gram atoms of SO <sub>2</sub> Contains:  (A) 6.02 x 10 <sup>23</sup> atoms of oxygen (D) 4 gram atoms of SO <sub>2</sub> How many particles are called fundamental particles of an atom?  (A) 3 (B) 5 (C) 100 (D) 6  What are the units of R <sub>1</sub> value?  (A) Cm (B) Cm <sup>3</sup> (C) dm <sup>3</sup> (D) No units  Which of the following cannot sublime?  (A) Naphthalene (B) locine (C) Ammonium chloride (D) MnO <sub>2</sub> if absolute temperature of a gas is doubled and the pressure is reduced to one half, the volume of the gas will  (A) Be doubled (B) Reduced 1/4 (C) Increases four times (D) Remain unchanged Partial pressure of oxygen in lungs (in) orr) is:  (A) 150 (B) 116 (C) 760 (D) 159  Molecules of CO <sub>2</sub> in dry ide form the:  (A) 150 (B) 116 (C) 760 (D) 159  Molecules of CO <sub>2</sub> in dry ide form the:  (A) 150 (B) 116 (C) 760 (D) 159  Molecules of CO <sub>2</sub> in dry ide form the:  (A) 150 (B) 160 (C) 170 (D) 2 positive rays  (B) 150 (D) 2 positive rays  (C) 200 n.m (D) 5 x 10 <sup>2</sup> m  Radioactive copper emits:  (A) 2 rays (B) β - rays (C) γ - rays (D) Positive rays  (B) 160 (C) 160 (D) 2 positive rays  (C) 2 positive rays  (D) 160 (D) 2 positive rays  (D) 2 positive rays  (D) 3 positive rays  (D) 3 positive rays  (D) 4 positive rays  (E) 10 positive rays  (E) 10 positive rays  (E) 10 positive rays  (E) 10 positive rays  (E)			D) Ba <sup>2+</sup> , NO <sub>2</sub> and Cl		
(A) 18 % solution of glucose (C) 5.85 % solution of sodium chloride (C) 5.85 % solution of sodium chloride (B) Fe Is precipitated out (C) Cu and Fe both dissolved (C) Cu and Fe both dissolved (D) No reaction takes place With increase of 10 °C temperature the rate of reaction doubles. This increase in rafe of reaction is due to: (A) Increase in number of effective collisions. (B) Increase in activation energy of reaction. (C) Decrease in activation energy of reaction. (D) Decrease in the number of collisions between reactant molecules.  One mole of SO <sub>2</sub> contains: (A) 6.02 x 10 <sup>23</sup> atoms of oxygen (C) 18.1 x 10 <sup>23</sup> molecules of SO <sub>2</sub> How many particles are called fundamental particles of an atom? (A) Cm (B) S (C) Idm (C) Idm (D) No units Which of the following cannot sublime? (A) Naphthalene (B) Iodine (C) Ammonium chloride (D) MnO <sub>2</sub> (D) Monopathalene (B) Iodine (C) Ammonium chloride (D) MnO <sub>2</sub> (D) Monopathalene (D) MnO <sub>2</sub> (D) Remain unchanged (D) Remain uncha				•	
(C) 5.85% solution of sodium chloride  If a strip of Cu metal is placed in a solution of FeSO <sub>4</sub> :  (A) Cu will be deposited  (B) Fe is precipitated out  (C) Cu and Fe both dissolved  (D) No reaction takes place  With increase of 10 °C temperature the rate of reaction doubles. This increase in rate  of reaction is due to:  (A) Increase in number of effective collisions.  (B) Increase in activation energy of reaction.  (D) Decrease in attivation energy of reaction.  (D) Decrease in the number of collisions between reactant molecules.  One mole of SO <sub>2</sub> contains:  (A) 6.02 x 10 <sup>23</sup> atoms of oxygen  (C) 18.1 x 10 <sup>23</sup> molecules of SO <sub>2</sub> How many particles are called fundamental particles of an atom?  (A) 3  (B) 5  (C) 100  (D) 4 gram atoms of SO <sub>2</sub> How many particles are called fundamental particles of an atom?  (A) 3  (B) 5  (C) 100  (D) 6  What are the units of R₁ value?  (A) Cm  (B) Cm <sup>3</sup> (C) dm <sup>3</sup> (D) No units  Which of the following cannot sublime?  (A) Naphthalene  (B) Iodine  (C) Ammonium chloride  (D) MnO <sub>2</sub> If absolute temperature of a gas is doubled and the pressure is reduced to one half , the volume of the gas will  (A) Be doubled  (B) Reduced 1/4  (C) Increases four times  (D) Remain unchanged  Partial pressure of oxygen in lungs (in forr) is:  (A) 150  (B) 166  (C) 200 in dry ice form the:  (A) Molecular crystals  (B) bonic crystals  (C) Covalent crystals  (D) Any type of crystals  Vapour pressure is not/affected by:  (A) Temperature  (B) Intermolecular forces  (C) Surface area  (D) Pressure  Wave number of the light emitted by a certain source is 2 x 10 <sup>6</sup> m <sup>1</sup> . The wavelength of this light will be:  (A) 500 n.m  (B) 500 m  (C) 200 n.m  (D) 5 x 10 <sup>7</sup> m  Sadioactive copper emits:  (A) A - rays  (B) β - rays  (C) γ - rays  (D) Positive rays  (D) Positive rays  (D) Positive rays  (E) A rays  (D) Positive rays  (E) A rays  (D) Positive rays  (E) A rays  (E)			(B) 60% solution of	urea	
If a strip of Cu metal is placed in a solution of FeSO <sub>4</sub> :  (A) Cu will be deposited  (B) Fe is preopitated out  (C) Cu and Fe both dissolved  (D) No reaction takes place  With increase of 10 °C temperature the rate of reaction doubles. This increase in rate  of reaction is due to:  (A) Increase in number of effective collisions.  (B) Increase in activation energy of reaction.  (D) Decrease in activation energy of reaction.  (D) Decrease in the number of collisions between reactant molecules.  One mole of SO <sub>2</sub> contains:  (A) 6.02 x 10 <sup>23</sup> atoms of oxygen  (C) 18.1 x 10 <sup>23</sup> molecules of SO <sub>2</sub> (D) 4 gram atoms of sulphur  (C) 18.1 x 10 <sup>23</sup> molecules of SO <sub>2</sub> (D) 4 gram atoms of SO <sub>2</sub>					
(A) Cu will be deposited (C) Cu and Fe both dissolved (D) No reaction takes place With increase of 10 °C temperature the rate of reaction doubles. This increase in rate of reaction is due to: (A) Increase in number of effective collisions. (B) Increase in activation energy of reactants. (C) Decrease in activation energy of reaction. (D) Decrease in the number of collisions between reactant molecules. One mole of SO <sub>2</sub> contains: (A) 6.02 x 10 <sup>23</sup> atoms of sulphur (C) 18.1 x 10 <sup>23</sup> molecules of SO <sub>2</sub> How many particles are called fundamental particles of an atom? (A) 3 (B) 5 (C) 100 (D) 4 gram atoms of SO <sub>2</sub> How many particles are called fundamental particles of an atom? (A) 3 (B) 5 (C) 100 (D) 6 What are the units of Rf value? (A) Cm (B) Cm <sup>3</sup> (C) dm <sup>3</sup> (C) dm <sup>3</sup> (D) No units Which of the following cannot sublime? (A) Naphthalene (B) Iodine (C) Ammonium chloride (D) MnO <sub>2</sub> If absolute temperature of a gas is doubled and the pressure is reduced to one half, the volume of the gas will (A) Be doubled (B) Reduced 1/4 (C) Increases four times (D) Remain unchanged 1 Partial pressure of oxygen in lungs (in torr) is: (A) 150 (B) 116 (C) 760 (D) 159 Molecules of CO <sub>2</sub> in dry ide form the: (A) Molecular crystals (B) Ionic crystals (C) Covalent crystals (D) Any type of crystals Vapour pressure is not affected by: (A) Temperature (B) Intermolecular forces (C) Surface area (D) Pressure Wave number of the light emitted by a certain source is 2 x 10 <sup>6</sup> m <sup>-1</sup> . The wavelength of this light will be: (A) 500 n.m (B) 500 m (C) 200 n.m (D) 5 x 10 <sup>7</sup> m 58 Radioactive copper emits: (A) α - rays (B) β - rays (C) γ - rays (D) Positive rays Which of the following molecules have zero dipole moment? (A) NH <sub>3</sub> (B) CHCl <sub>3</sub> (C) BF <sub>3</sub> (D) H <sub>2</sub> O The bond order of helium molecule is: (A) 3 (B) 2 (C) 1 (D) Zero			1 7 1 1/	Zeimig Permi	
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(A) Cm (B) Cm³ (C) dm³ (D) No units  Which of the following cannot sublime?  (A) Naphthalene (B) Iodine (C) Ammonium chloride (D) MnO2  If absolute temperature of a gas is doubled and the pressure is reduced to one half, the volume of the gas will  (A) Be doubled (B) Reduced 1/4 (C) Increases four times (D) Remain unchanged Partial pressure of oxygen in lungs (in torr) is:  (A) 150 (B) 116 (C) 760 (D) 159  Molecules of CO2 in dry ide form the:  (A) Molecular crystals (B) onic crystals (C) Covalent crystals (D) Any type of crystals Vapour pressure is not affected by:  (A) Temperature (B) intermolecular forces (C) Surface area (D) Pressure Wave number of the light emitted by a certain source is 2 x 10 <sup>6</sup> m <sup>-1</sup> . The wavelength of this light will be:  (A) 500 n.m (B) 500 m (C) 200 n.m (D) 5 x 10 <sup>7</sup> m  Radioactive copper emits:  (A) α - rays (B) β - rays (C) γ - rays (D) Positive rays  Which of the following molecules have zero dipole moment?  (A) NH <sub>3</sub> (B) CHCl <sub>3</sub> (C) BF <sub>3</sub> (D) H <sub>2</sub> O  The bond order of helium molecule is:  (A) 3 (B) 2 (C) 1 (D) Zero		/			
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	1/ (OD)/1 A		7.		

**CHEMISTRY** TIME: 2 HRS 40 MINUTES **GROUP: FIRST MARKS: 68** SUBJECTIVE PART SECTION-I DGK-1-24 16 QUESTION NO. 2 Write short answers to any Eight (8) of the following N<sub>2</sub> and CO have the same number of electrons, protons and neutrons, justify. Law of conservation of mass have to be obeyed during stoichiometric calculations, explain. ij iii Why actual yield is always less than theoretical yield? iv Write two suitable uses of the technique of chromatography In solvent extraction technique, why repeated extractions using small portions of solvent are more efficient than using a single extraction but larger volume of solvent. How undesirable colours in crystallization process can be removed? νi Write formulas to interconvert various scales of temperature. vii How density of an ideal gas can be calculated from ideal gas equation? viii Derive Charle's law by kinetic equation of gases. ix What is Handerson equation and for what purpose it is used? X What are applications of buffer solutions in daily life? xi Derive ionic product of water and what is its value at 25°C Xii

QUES	TION NO. 3 Write short answers to any Eight (8) of the following	16
i	Why intermolecular forces are weaker than intramolecular forces?	
ii	What are advantages of Vacuum distillation ?	
iii	Differentiate between Isomorphism and polymorphism.	
iv	Diamond is hard and electrical insulator. Justify it.	
v	Explain Atomic Emission Spectrum.	
vi	Define (a) Wave number (b) Frequency	
vii	Write electronic configuration of Cr <sub>24</sub> and Zn <sub>30</sub>	
viii	What is Moseley's law ? Give its mathematical expression.	
ix	What do you mean by water of crystallization? Give an example.	
х	Why NaCl and KNO <sub>3</sub> are used to lower the melting point of ice?	
xi	Differentiate between instantaneous and average rate of a reaction.	
xii	What do you mean by Homogeneous catalysis ? Give an example.	
	and the same of th	

QUES	STION NO. 4 Write short answers to any Six (6) of the following	12
i	How does the hybridization scheme explain the bond length?	
ii	Define electron affinity. Name the factors affecting it.	
iii	The radius of an atom cannot be determined precisely. Give the reason.	
iv	Why do the lone pairs of electrons on an atom occupy more space than bond pairs?	
v	Define standard enthalpy of formation. Give an example.	
vi	Define exothermic reaction. Give an example.	
vii	Differentiate between spontaneous and non-spontaneous process.	
viii	What is anodized aluminium?	
iv	Give the electrode reactions during the recharging of lead accumulator.	

## SECTION-II

Note: A	ttempt any Three questions from this section	$8 \times 3 = 24$
Q.5.(A)	Define limiting reactant, write down the steps involved in identification of limiting reactant.	1+3
(B)	Define hydrogen bonding , how does it explain structure of ice (without diagram).	1+3
Q.6.(A)	Write a note on " Principal Quantum Number"	4
(B)	250 Cm <sup>3</sup> of the sample of hydrogen gas effuses four times as rapidly as 250 Cm <sup>3</sup> of an unknown gas. Calculate the molar mass of unknown gas.	4
Q.7.(A)	Discuss sp – hybridization with example of ethyne.	1+3
(B)	The solubility product of Ag <sub>2</sub> CrO <sub>4</sub> is 2.6 x 10 <sup>-2</sup> at 25°C. Calculate the solubility of the compound	d. 4
Q.8.(A)	Describe construction and working of a Bomb Calorimeter.	4
(B)	What is standard electrode potential? How can it be measured?	4
Q.9.(A)	What are continuous and discontinuous solubility curves? Draw these curves to explain the answer	er. 2+2
(B)	Discuss homogeneous and heterogeneous catalysis in detail with two examples of each.	2+2