

Roll No. of Candidate : _____

CHEMISTRY

(INTERMEDIATE PART-II) 421 - (II) Paper II

(Group - I)

Time: 20 Minutes

OBJECTIVE Code: 8483

Marks: 17

Note: You have four choices for each objective type question as A, B, C and D. The choice which you think is correct, fill that circle in front of that question number. Use marker or pen to fill the circles. Cutting or filling two or more circles will result in zero mark in that question. Attempt as many questions as given in objective type question paper and leave others blank.

60J-51-21

1. The strength of binding energy of transition elements depends upon:
(A) number of electron pairs (B) number of unpaired electrons
(C) number of neutrons (D) number of protons
2. Which compound shows maximum hydrogen bonding with water?
(A) CH_3OH (B) $\text{C}_2\text{H}_5\text{Cl}$ (C) $\text{CH}_3-\text{O}-\text{CH}_3$ (D) $\text{C}_6\text{H}_5\text{OH}$
3. SN_2 reactions can be best carried out with
(A) primary alkyl halides (B) secondary alkyl halides
(C) tertiary alkyl halides (D) all of these
4. The solution of which acid is used for seasoning of food?
(A) formic acid (B) acetic acid (C) benzoic acid (D) butanoic acid
5. The oxide of Beryllium is
(A) acidic (B) basic (C) amphoteric (D) none of these
6. Which of the following compounds will not give iodoform test on treatment with I_2 / NaOH ?
(A) acetaldehyde (B) acetone (C) butanone (D) 3-pentanone
7. Which of the following is not a fatty acid?
(A) propanoic acid (B) acetic acid (C) phthalic acid (D) butanoic acid
8. Which is the strongest acid?
(A) HClO (B) HClO_2 (C) HClO_3 (D) HClO_4
9. Which element belongs to group IVA of the periodic table?
(A) barium (B) iodine (C) lead (D) oxygen
10. Micro-nutrients are required in quantity ranging from
(A) 4 - 40 g (B) 6 - 200 g (C) 6 - 200 kg (D) 4 - 40 kg
11. Select from the following the one which is alcohol
(A) $\text{CH}_3-\text{CH}_2-\text{OH}$ (B) $\text{CH}_3-\text{O}-\text{CH}_3$ (C) CH_3COOH (D) $\text{CH}_3-\text{CH}_2-\text{Br}$
12. Which of the following species has the maximum number of unpaired electrons?
(A) O_2 (B) O_2^+ (C) O_2^- (D) O_2^{2-}
13. Which of the following has the lowest melting point?
(A) Be (B) Mg (C) Ca (D) Sr
14. Which one of the following compounds will react with Fehling's solution?
(A) $\text{C}_2\text{H}_5\text{COOH}$ (B) CH_3CHO (C) CH_3COOH (D) CH_3COCH_3
15. β - β' - dichloroethyl sulphide is commonly known as
(A) laughing gas (B) mustard gas (C) phosgene gas (D) bio-gas
16. The benzene molecule contains
(A) three double bonds (B) two double bonds
(C) one double bond (D) delocalized π -electron charge
17. The halogen with the lowest melting and boiling points is
(A) fluorine (B) chlorine (C) bromine (D) iodine

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Time: 2:40 Hours

SUBJECTIVE

Marks: 68

Note: Section I is compulsory. Attempt any THREE (3) questions from Section II.(SECTION - I) 40J-41-21**2. Write short answers to any EIGHT questions.**

(2 x 8 = 16)

- Why the size of an anion is larger than its neutral atom?
- What is the role of shielding effect on ionization energy?
- Write down electronic configuration of Na and Ca.
- Why the group I-A elements are called alkali metals?
- Give four uses of borax.
- Write down two points about the peculiar behaviour of carbon from its group.
- What happens when borax is heated with NH_4Cl . Write down balanced equation.
- Write down formulas of the following minerals:
(a) Galena (b) Heavy Spar
- Sulphuric acid is a dehydrating agent. Prove it by giving two equations.
- Briefly describe the role of nitrogen in plants.
- Write down the major steps involved in the synthesis of urea fertilizer.
- What are the raw materials used in the manufacture of cement?

3. Write short answers to any EIGHT questions.

(2 x 8 = 16)

- Prepare Cl_2O_7 with the help of chemical reaction.
- Prepare HClO_4 . Also write down its two properties.
- Write down any four uses of bleaching powder.
- Write down the name of any four methods for prevention of corrosion.
- How Zinc coating or anode coating prevents the iron from corrosion?
- Describe the x-rays structure of benzene.
- Prepare benzene and toluene from alkane with equation.
- Write down the reaction of acetone with 2, 4 - dinitrophenylhydrazine.
- Write down any four uses of acetaldehyde.
- How acetic acid is prepared from acetylene?
- Write down the chemical reaction of CH_3COOH with (i) $\text{C}_2\text{H}_5\text{OH}$ (ii) NH_3
- How would you convert acetic acid into acetic anhydride?

4. Write short answers to any SIX questions.

(2 x 6 = 12)

- Define geometric isomerism giving one example.
- What are aliphatic compounds? Give their two examples.
- What is clemmensen reduction? Give example.
- Convert (a) Methane into ethane (b) Ethene into ethylene glycol.
- State Markownikoff's Rule with an example.
- Define nucleophile and substrate. Giving one example in each case.
- Convert ethyl chloride into (a) Ethane (b) Tetraethyl Lead
- What is denaturing of alcohol?
- How will you distinguish between ethanol and methanol by a chemical test?

(SECTION - II)

- (a) Define electron affinity. How does it vary in groups and periods generally in the periodic table. 4
(b) Give the formula of Sylvite, Borax, Trona, Natron, Dolomite, Alunite, Asbestos and Barite. 4
- (a) Discuss the preparation of nitric acid by Birkeland and Eyde's process. 4
(b) Discuss the binding energies and oxidation states of transition elements. 4
- (a) Write down note on reforming of gasoline. 4
(b) Explain oxidation of aldehydes and ketones with two examples in each case. 4
- (a) How alkanes can be prepared by Kolbe's electrolytic method. Write down its mechanism. 4
(b) What is β -Elimination reaction? Differentiate between E_1 and E_2 elimination reactions. 4
- (a) Describe Kekule's structure of benzene. 4
(b) How does ethanol react with 4
i) Na ii) PCl_5 iii) CH_3MgI iv) SOCl_2 4

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